

Mobile phase considerations

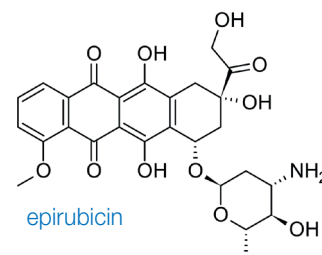
In HILIC the mobile phase contains high concentrations of organic solvent (60–9%) and low concentrations of water (3–40%). This provides the advantage of a lower viscosity and reduced back pressures, even with relatively long columns. The high volatility of the eluent leads to increases in sensitivity of evaporative detectors.^[7] However the different overlapping interaction mechanisms in HILIC and especially the formation of a stable water layer on the stationary phase surface also leads to longer equilibration times, which is even more important when using gradient methods.

Acetonitrile (AcN) is the most widely used organic solvent in analytical HILIC, due to its good UV transparency and low viscosity. When increasing the AcN content in the mobile phase and reducing the water concentration, retention significantly increases, but the increase in retention factor doesn't follow a linear trend.^[2]

This is due to the overlap of adsorption and partition mechanisms in HILIC separations and the change in the thermodynamic pH, which can occur when the concentration of an organic compound in the mobile phase changes. Full or partial substitution of AcN with alcohols such as methanol (MeOH), ethanol (EtOH) or isopropanol (IPA) will reduce retention by weakening hydrogen-bonding and disrupting the formation of the stationary water layer, which usually has a negative effect on reproducibility.^[8-12]

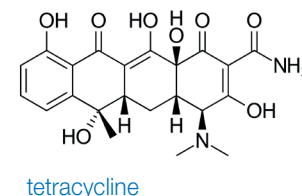
A ranking of solvents in order of relative solvent strength is not easily estimated in HILIC because of the different overlapping separation mechanisms and strongly depends on the choice of stationary phase. For instance, separation of polar analogues of the inhibitory drug epirubicin on bare silica gives a ranking in the order of:

methanol < isopropanol < tetrahydrofuran < ACN



whereas polar tetracycline antibiotics on an amino-bonded phase show the ranking in order of:

tetrahydrofuran < methanol < isopropanol < ACN

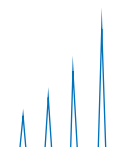


Notably AcN is the weakest solvent regardless of testing conditions.^[13]

Retention of weaker acids and bases is greatly affected by mobile phase pH, while strong acids and bases are affected to a lesser extent. The main retention mechanism for these compounds are coulombic interactions between the stationary phase charged groups and the charged analytes. The influence of pH on the charged state of the stationary phase is also a big factor for these compounds.

Small acidic mobile phase modifiers such as **formic acid**, **acetic acid** or phosphoric acid can act as reasonable “buffers” or at least keep a constant pH, but generally do not have adequate ionic strength in high AcN concentrations to reduce

silanol activity and produce symmetric peak shapes.^[14-15] Because silanols on the stationary phase surface will be fully protonated at pH 2–3, TFA or HFBA is often used to reduce the net charge of residual silanols to zero and to suppress the cationic exchange mechanism with basic analytes.^[16] However this can produce an effect of increased repulsion and even exclusion of acidic analytes from the stationary phase especially on **bare silica** and **amide** phases. It is believed this may be due to the enrichment of hydronium ions on the water/eluent interface and/or charged AcN artifacts due to hydrolysis but these effects still require further investigation.^[17-19]



The use of **salts derived from formic acid and phosphoric acid, as well as oxalic acid and citric acid as buffer ions** is a well-established tool to alter retention times and improve peak shapes, but this shows little effect on the selectivity of the separation itself.^[15,20] This means that higher resolution and sensitivity can be achieved by screening a method for the appropriate buffer concentration and its components. The most widely used buffer systems in HILIC are **ammonium formate and -acetate** because of their solubility in high AcN concentrations and also their volatility, which makes them a suited buffer system for use in ESI-MS.

The use of buffer systems of pH9 or higher can significantly change selectivity, especially as quaternary ammonium compounds show high retention under these conditions due to ionic interactions with deprotonated silanols. Since most silica-based materials will rapidly dissolve under

these conditions, the use of **hybrid-silica** phases is a novel tool to overcome this obstacle and make it possible to screen a high pH values in method development.

The **concentration of the buffer** used has a large effect on retention of analytes using electrostatic interactions as the main separation mechanism. Variation in the range of 5–20mM buffer concentration can lead to a drastic change of selectivity.^[2] Combined with the nature of (high organic) liquids to evaporate over time this is a common source for non-reproducible results when using eluents which have not been freshly prepared. On the other hand this “shielding” effect of silanols due to counter ions in the mobile phase can lead to useful selectivity effects if buffer concentration can be maintained reproducibly. Furthermore increased salt concentration also increases the volume of the adsorbed water layer which successively increases the retention of neutral solutes.^[21-22]

Column temperature

An increase in column temperature is always accompanied by a reduction in backpressure due to the decrease in viscosity of the mobile phase. But in HILIC mode temperature has a complex effect on the ionisation of stationary phase bonded groups, silanols, buffer components and analytes. Variation in the temperature range of 30–50°C generally shows a decrease in analyte retention of 3–30% with increasing temperature.^[2, 12] However some basic analytes

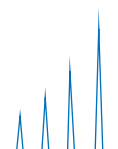
show increased retention with increasing temperature, which is an effect that currently cannot be explained. Due to the relatively small effect of temperature changes on selectivity, it is often an overlooked parameter in method development. To ensure reproducible results column temperature should at least be set to a value just over room temperature to compensate for possible temperature fluctuations in the environment.

Sample injection

The influence of the injection solvent on HILIC separations and the resulting peak shapes is well studied on small molecules as well as peptides.^[23] Generally samples should be injected in **pure AcN** to give best peak shapes and resolution. **Higher water concentrations in the injection solvent will proportionally lead to loss of efficiency** due to a plug of strong solvent and analytes being eluted through the column.^[20]

In cases where it is not possible to inject in AcN because of the solubility of sample constituents or aggregation of peptides, it has been shown that reduction in injection volume can minimise this effect. As an example protein and antibody biopharmaceutical analysis

in HILIC gave satisfactory results at concentrations of 65–80% AcN when injecting in pure water, but decreasing injection volume to 0.1% of the column volume.^[24] This strategy might also be applicable to environmental samples from pure water, but the generally low analyte concentrations in these applications may require very sensitive detection methods suited for HILIC or an enrichment of analyte via sample preparation techniques. For small molecules which are insoluble in AcN, it may be possible to inject in IPA or IPA/AcN mixtures providing solubility can be assured. For peptide analysis a switch to pure organic solvents such as EtOH or IPA can limit denaturation from occurring.^[23]



Detection procedures

Because high concentrations of volatile solvents are used in HILIC, the response of different detection methods in which the mobile phase is evaporated is greatly enhanced.^[25]

Furthermore the formation of charged droplets in CAD or ESI-MS detection is facilitated by the low surface tension of HILIC mobile phases. The design of the ESI source has a huge influence on the relative sensitivity, but for most analytes an increase of signal-to-noise ratio compared to RP

can be observed in the range of one order of magnitude and can increase by up to a factor of 800.^[26]

It is very important particularly for methods using these detectors to maintain a constant pH and buffer concentration. When working with ESI-MS screening at different buffer concentrations can help to evaluate the ion suppression effects of buffer salts which can counteract some of the sensitivity gained.

Conclusions and summary

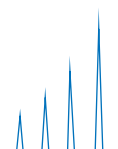
As an orthogonal separation mode HILIC offers unique advantages over reversed phase chromatography due to the high organic eluent contents and different interaction mechanisms. As a result of the water layer on the stationary phase, mass transfer and analyte diffusivity play a greater role in HILIC kinetics. Consequently HILIC is less well understood than other LC modes.

The greatest changes in selectivity are achieved by using different stationary phase chemistries, while changing organic solvent concentration and running gradients usually

has a bigger influence on retention than selectivity itself. pH and the use of different buffer salts and concentrations have a considerable effect on analyte charge, stationary phase charge, water layer volume and selectivity of the stationary phase due to silanol activity quenching. Temperature has little effect on selectivity, but it is advised to keep it constant to maintain reproducibility.

In summary, we can give the following general tips for reproducible HILIC method development and troubleshooting:

- Stationary phase selectivities are very different in HILIC analysis. Screening different phases may give you a more appropriate phase for your analytes.
- The mobile phase should contain at least 3% and a maximum of 40% water to produce a stable water layer on the stationary phase.
- Depending on the application we suggest buffer concentrations up to 10 mM and to buffer both mobile phases for better miscibility and reproducibility
- Recommended buffers are ammonium salts of acetic or formic acid, bicarbonate salts or triethylamine phosphate for high solubility in organic solvents and suppression of secondary interactions. Acids such as formic acid or TFA as modifier may give poorer results than the use of genuine buffers.
- Use aprotic solvents such as acetonitrile, THF or acetone as weak eluent. Use of protic solvents including alcohols generally decreases retention and disturbs the stationary phase water layer.
- Keep in mind the pKa of your analytes and their charged state at a given pH. It is good practice to work 1–2 pH units above or below the pKa of your analytes and screen using at least 3 different pH values in your method development.
- Dissolve your sample in high organic concentration (pure AcN or other aprotic or protic solvents) or up to the starting composition of your eluent to suppress dilution and negative partition effects. Never inject samples in higher content of water or if this is not possible reduce the injection volume significantly.
- Give your HILIC phase sufficient time for equilibration. We recommend at least 20 column volumes prior to analysing samples and/or after gradient elution. If running a gradient keep the gradient slope relatively low or run small plateaux.



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